

ACETYLENE

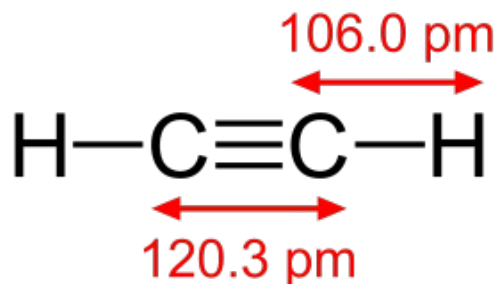


by Miles Mathis

First published February 2, 2024

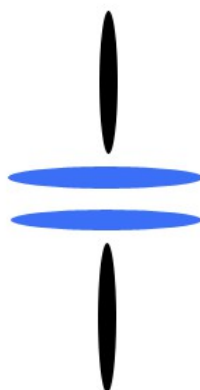
[This paper immediately went to #5 at Yandex on the search “acetylene diagram”. Google, Bing and Yahoo are illegally downlisting it by instead multi-listing and looping NIST Chemistry WebBook, Chemistry Libretexts, byjus.com and others.]

I just did Ammonia, and Acetylene is used in the production of Ammonia, so let's do it next. It is a good way to prove again my nuclear diagramming models.

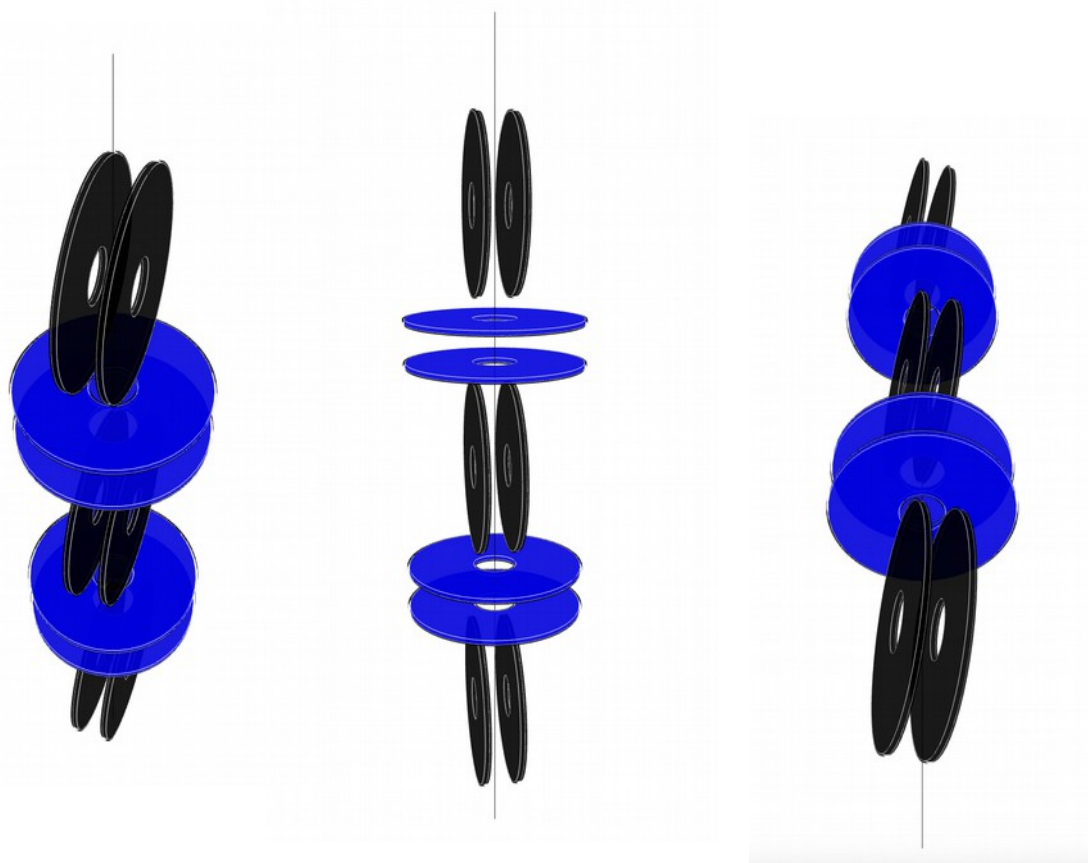


Those are the mainstream diagrams of Acetylene, which show Carbon backed in with itself in the central bond. Carbon is not a gas or diatomic, so that configuration doesn't make much sense given mainstream theory. Carbon shouldn't want to bond with itself, since if it did it would in normal circumstance. Yes, Dicarbon exists, but it is an unstable vapor that quickly reacts to form other compounds. But using my diagram of the Carbon nucleus explains Acetylene immediately:

Carbon



The blue alphas—being made of two protons—can accept two protons in the N/S plugs (in some circumstances), so Carbon has free slots north and south. Acetylene simply fills all available slots.



illustrations by Arlo Emerson

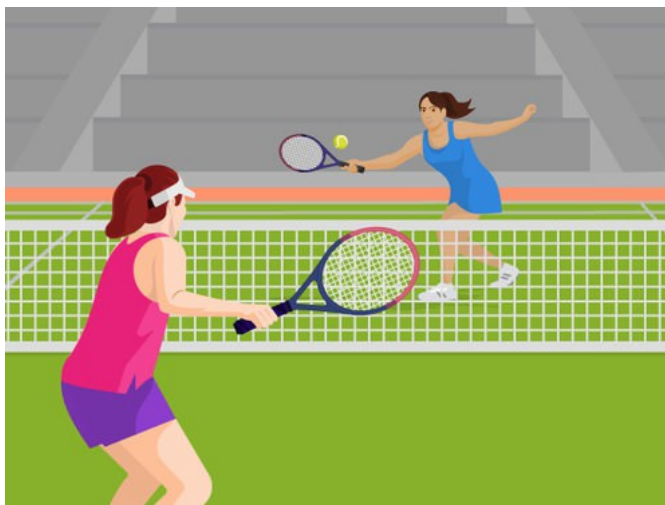
So, other than that triple bond they claim, my diagram matches the mainstream here, being linear and symmetrical. But they don't realize this is linear along the Carbon nuclear poles, and that it is not fully symmetrical on Earth due to imbalance in the Terrestrial and Solar field. Charge will always be moving through this molecule more strongly in one direction than the other.

Besides, Acetylene isn't some new element, a variant form of Silicon, with six protons plugged in tightly. The two Hydrogens are backed-out some distance in molecular positions, not in nuclear positions. And the bond between Carbons is molecular as well, not a fused double nucleus. So the alphas aren't really channeling two parcels of charge.

But what does the mainstream mean by the triple bond between the Carbons? We are told that involves six electrons from the two Carbons. But Carbon allegedly has four electrons in the outer shell, two of them *s* and two *p*. The *2s* level is already full, so I don't understand why it would want more. At any rate, you should ask why Carbon would bond to itself with 3 of those valence electrons, and not 2 or 4? The historical answer is they were trying to match a measured bond strength, which was not what they expected from their *s* and *p* orbitals. So they created a very messy fudge with a “hybridized” *sp* orbital, which combines **all four** *s* orbitals in **one** big sigma bond. The *p* electrons also then bond, so all 2nd shell electrons are used. My readers will say, “Isn't that just something they pulled out of their shorts?”

Yep. They have no evidence that is what is happening, other than the measured bond strength. This is the way they matched it, so it must be true, they say.

The other bonds in Acetylene are covalent, which brings us back to that catastrophe. In doing a new general search on the covalent bond, the first item Bing returned was [this link to chem.libretexts.org](http://chem.libretexts.org). There we find them leading with this:



In a tennis match, two players keep hitting the ball back and forth. The ball bounces from one player to the other, over and over again. The ball keeps the players moving together on the court. What if the two players represented the nuclei of two atoms and the ball represented valence electrons? What would the back and forth movement of the ball represent? The answer is a covalent bond.

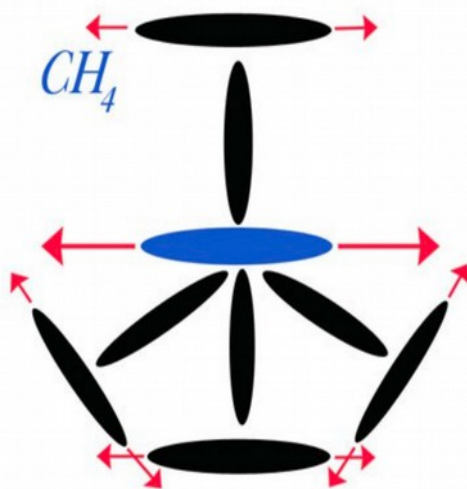
Wow. So this is what we are teaching our young people? “The ball keeps the players moving together on the court”. As a matter of physics, chemistry, or anything else, is that true? No, it is a bald logical fallacy, since the ball isn't causing anything, much less linking the two players. The players are reacting to one another, not the ball, and they could just as easily not. This is a near-perfect example of anti-science. Nothing even vaguely resembling a molecular bond is going on here, so this is just bald propaganda for the standard model, brainwashing kids with obvious absurdities. The usual.

But the mainstream really doesn't have anything better than this to tell students after a century of working on the problem. You may think this is just silly and pushed analogy for five year olds, but it is state of the art in chemistry. They have been telling us all along that a real bond can be created by crossing paths. Or something. No one has ever made sense of it, or even seriously tried.

Of course I have a much cleaner explanation of the triple bond. Electrons have nothing to do with it and this is just a charge bond, with those two protons acting as fans between the Carbons. So I would draw it as a double proton bond, not a triple electron bond. The reason it is stronger than you might think and stronger than they predicted doesn't have anything to do with hybridized orbitals or fudged equations, it has to do with several factors, some of which I have pointed out in previous papers. When an element bonds to itself like this, the charge streams already match in strength. They don't have to be stepped up or down to match, so they are at maximum efficiency. So they appear stronger than the same bonds between *different* elements. Acetylene also benefits from a fairly rare configuration that you can see is both linear and symmetrical, which also maximizes charge channeling. This problem

could have been avoided if bond strengths had been pegged from the beginning to maximum values like this: linear, symmetrical, and between the same element. Then other similar bonds of the same type would be a percentage of that, and we would calculate losses due to angles, charge density changes, and lack of field symmetry. As it is, these equations have to be heavily fudged in the other direction (up), with no possible mechanics to point to. Just a lot of unverifiable—and bad—guesses that orbitals are being hybridized in some squishy way to match numbers.

So let's back up: according to my theory, why does Carbon bond with itself in Acetylene, but not normally in a diatom. The mainstream glosses right over that question, but again I can answer it easily. It is because the CH forms first, which then facilitates the central CC bond. Carbyne is very common, though it doesn't persist. But Acetylene is commonly made by combustion of Methane. The north protons are blown off Methane by Oxygen, leaving a lot of highly reactive Carbyne floating around. Some of it then reforms into Acetylene. Why? The mainstream tells us it is because Carbyne is a radical, having one or three unpaired electrons, but that is the usual fudge. There is no evidence for that because it isn't happening. What is happening is that in the combustion of Methane, Carbyne has been left with two protons in the south hole and none in the north, giving it a double charge imbalance. Let's go back to my diagram of Methane so you can see this:



We start with that, though it is upside down. The big part of the tetrahedron is the weak end, so it should be in the north position, to match the profile of the Earth. I originally matched it to the mainstream diagram, which has the small end north simply because that is how you normally draw a tetrahedron. Anyway, in the presence of Oxygen at the right energy, the larger charge streams created by the nucleus of Oxygen send too much charge through Methane, blowing the three cap protons (Hydrogens) off the bottom end of that diagram. Charge is coming *out* down there, so the elevated charge stream will blow all that away, while leaving the top end unaffected. Charge is coming in up there, moving straight down, so if anything it will press the top in even more. But we do find one other change. That cap proton on top will be blown parallel with the one just beneath it, given us two plugged into the blue alpha. Which gives us Carbyne. You can see the Carbyne is now very top-heavy, with two protons in the top pole there, and none in the bottom. This gives it what the mainstream calls an affinity. It wants to fill those bottom holes, *but it isn't an electron affinity*. It is a proton affinity, and it wants them because it wants to create balance. The charge streams on the Earth are not balanced, but they are bi-directional, moving through all atoms and molecules in both

directions: that is, south-to-north and north-to-south. The first normally dominates 2-1 on the Earth, but nuclei and molecules maintain balance as much as they can in this bi-polar field, since balance is the way to avoid destruction. It is the way to avoid radioactivity and dissolution. This is the charge pressure that keeps them together from both ends, and without they will decay in some way. Molecules will decay down to nuclei.

You would think Carbyne would react first with Hydrogen, but in this manufacture by combustion, the Hydrogen is drawn more strongly to Oxygen, creating water. Other methods of dehydrogenation are used to keep Carbyne from reacting with Hydrogen, leaving it to react with itself. In that situation, it wants to do that far more than Carbon wants to bond with itself. Why? Because that extra Hydrogen on the pole has not only increased its charge channel, it has focused it. We saw a similar thing with [Fluorine plugged in the pole of larger molecules](#), making them much more reactive. The same thing is happening here, but on a smaller scale. In fact, Carbyne has a similar structure to Fluorine, with two protons plugged in the primary pole. But both of Fluorine's two protons in that position are nuclear, while one of Carbyne's is just molecular—backed-out some distance. So while Carbon by itself isn't very conducive to a diatom—due to a weakly directionalized charge field—the extra proton in the primary pole makes Carbyne much more reactive. So we don't need electrons to explain anything here. We have a much better explanation of all these phenomena without electrons.