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# The Great Methane Stink



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To start this paper, we will look at [a webpage](#) on Methane created by a United Kingdom textbook author from Cambridge. I don't mean to pick on him; I link to him only for convenience sake. He is no worse than anyone else, and is only teaching mainstream theory. I usually start at Wikipedia in my analysis of mainstream theory, since so many people go there to get basic questions answered; but Wiki has nothing for us on this topic. So to help my readers start at ground level, I searched on “why is Methane diagrammed as a tetrahedron?” That might be the first question one would ask, supposing one were not prone to accept uncritically everything one were told. Jim Clark's site is the first site that comes up on that question, and I found it to be pretty transparent.

Clark admits very quickly that there is “a mismatch” between data and theory regarding Methane. If we try to create the bonds using electron orbital theory, we find Carbon only has two unpaired electrons. Clark tells us,

The only electrons directly available for sharing are the 2p electrons. Why then isn't methane CH<sub>2</sub>?

That's a great question. But an even better one is already buried: how can sharing electrons create a bond? Is the bond supposed to be created by overlapping paths, and if so, how? Or is the bond created by overlapping probabilities? How, exactly, does superimposed math create a physical bond? Further, how can elements share electrons? Further, why would elements that *aren't ionized* wish to share electrons? Aren't they *already* stable? Why and how would an element become unstable in order to create a bond that re-created stability? Isn't that illogical?

The mainstream doesn't answer any of those questions, it simply dodges them. The only question Clark answers is how the orbitals can be pushed to create four unpaired electrons. He says,

There is only a small energy gap between the 2s and 2p orbitals, and so it pays the carbon to provide a small amount of energy to promote an electron from the 2s to the empty 2p to give 4 unpaired electrons. The extra energy released when the bonds form more than compensates for the initial input.

Yes, but “compensating” after the fact isn't the question, is it? The question is, where does the energy

come from to promote the electron up in the first place? How does the Carbon atom know this needs to happen, and how does it make it happen? Does it borrow energy from the vacuum? Does it wait for a passing photon? Notice, to promote the electron to a higher energy, the Carbon atom has to borrow just the right amount from somewhere. If it borrows too little, the electron doesn't get promoted; too much and the electron escapes. To sell this theory of promotion requires you to believe that the Carbon is both self-monitoring and self-correcting, based on no stated field mechanics.

But it gets worse. Even that one push won't get us where we need to go. Clark admits this problem, too:

In methane all the carbon-hydrogen bonds are identical, but our electrons are in two different kinds of orbitals. You aren't going to get four identical bonds unless you start from four identical orbitals.

He then gives you the second push. We need the electrons to be identical, so mainstream theory simply *makes* them identical! In a process called “hybridization,” the four electron orbitals arrange themselves in space so that they are as far away as possible. This creates a tetrahedron.



So that is where the tetrahedron comes from. But that begs a passel of questions, starting with “how does the atom keep the electrons in those orbitals? What is to prevent each electron from going elsewhere?” Even if we define those shapes as probabilities, the nucleus must have some way of physically creating those maxima. Current theory never provides any mechanics to do that. It is theory by fiat. We needed it, therefore it must be! Hybridization isn't really a process, it is a hamhanded push to match data.

Here's another question you might ask: have the protons in the nucleus re-arranged to cause that new shape? If not, we should have a charge separation problem. Those new orbitals—even treated as probabilities—don't maintain the original charge separation with the nucleus. You can't change the wavefunctions of the electrons while ignoring the charge profile of the nucleus. But if the proton configuration has changed, then how has it changed and how did it change? More actions upon itself?

We also have no mechanics for the hybridization. How do the electrons equalize? How do they communicate across empty space, in order to hybridize? Quantum mechanics is supposed to have bypassed the old force-at-a-distance problems of Newton, but we see again they bypass the old problems only by ignoring them. No, it is more than an ignoring, it is a forbidding. They now forbid you from asking these questions, and attempt to shame you if you do. They say something like, “You obviously haven't studied quantum mechanics at the graduate level. If you had you would know that classical mechanics doesn't work at the quantum level. We are past all that.” But the truth is, they aren't “past all that.” They haven't progressed to a subtler understanding of physics, they have simply stopped trying to do mechanics. Then they boldly try to sell you this failure as a success. “We couldn't solve the problems mechanically, but we are better for it. Mechanics is low and mundane, and we are ascended and esoteric. If you are still expecting physics to make sense, you are just a throwback. Go back to your abacus and your sundial!” The new science is just a psy-op, like everything else.

Next, Clark tells us how the bond is formed. According to mainstream theory, the orbitals “merge.” To illustrate this, Clark just imports the entire hydrogen atom into the green orbital (see last illustration). The green orbital was already a hybridized s-p orbital, so the new merged orbital is a 2s-2p-1s orbital, I guess. Don't ask how any of this works, since they can't tell you. Shouldn't the proton in the Hydrogen reject the extra electron now flying around it? If Hydrogen could so easily accept new electrons, why don't we see free Hydrogen negatively ionized like this? Remember, Hydrogen *approached* this bond, for some reason, although it already had its own electron. Why would it do that? According to the basic principles of E/M, it should flee this bond. To answer that, you are shown a number for electronegativity or electron affinity, but that is not an answer. Again, those numbers comes *after the fact*, and are just an indication that the bond is possible. But we already knew that from experiment. What we want from theory is an explanation of the mechanics. [The electron orbital theory doesn't provide any explanation.](#) It just pushes orbitals that don't work into ones that do, by main force.

Another way this flouts basic E/M theory founded on opposite charges is that the energy levels of those shared electrons is still not equal in any way: neither in shape nor energy nor anything else. This causes at least two major problems, both of which should be fatal. We are told that Carbon is now sharing half of eight electrons, which keeps the negative charge in the 2-level at 4. But since Carbon is sharing half of the Hydrogens' four 1s electrons, the four merged orbitals should collapse in some way. After the bond, Carbon has lost half of a 2s-2p orbital and gained half an 1s orbital. This should create a wavefunction compression. But are there any rules to such compressions, or can molecules adjust to accommodate any energy levels we throw at them? We have just seen the Carbon atom adjusting to several *ad hoc* combinations to suit these theorists. But if the Carbon nucleus is so accommodating, why do we see the definite orbitals to start with? We are taught that Carbon's original 2-level was created to keep it from interfering with the 1-level. But now we see Carbon's 2-level collapsing to accommodate a double-hybridized s-p-s merged orbital. Why doesn't that merged orbital interfere with Carbon's 1-level electrons? Apparently, these orbitals are set in stone. . . until we wish to move them, in which case they go wherever we want them to.

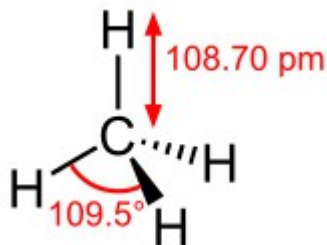
The second problem is the exclusion of electrons, which is simply ignored here. If electrons are negatively charged, they should exclude one another. In other problems, this is admitted and exploited, but in electron sharing, it is inverted. In this problem, electrons that don't even match orbitals or energy levels are jammed together to suit theorists, and not only is their exclusion turned off, but their proximity is actually used to create the bond. Remember, a bond is not an exclusion, it is an *attraction*. Sharing electrons creates an attraction.

To see what I mean, suppose I proposed that protons—which normally exclude one another—suddenly created a bond when they were shared. You would likely say I was mad and demand from me a mechanics that would explain that. And yet you demand no mechanics from [the madmen who have sold you the theory of electron sharing.](#)

Notice that I could explain the strong force as this sharing of protons. All I need to do is follow the logic of electron sharing. Normally, protons exclude one another strongly. But in the nucleus, the protons *are shared*. When they are shared, they attract. Problem solved.

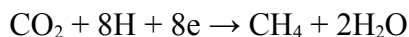
I am NOT proposing that, so don't worry. I have already gotten rid of the strong force by more logical means. We know from data that protons and electrons do exclude one another, which should have logically forbidden any theories of sharing a century ago. Methane can't be created this way, so we

know that it isn't.

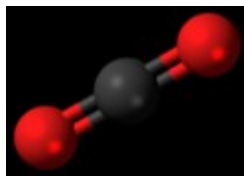


But we know the angle of Methane is 109.5. Doesn't this confirm the theory? No. It may confirm the tetrahedron, but it doesn't even begin to confirm electron orbital theory. To prove that, all I have to do is show an easier way to produce the tetrahedron, while avoiding any mention of electrons.

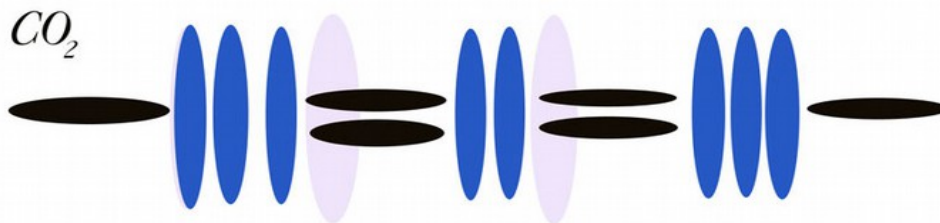
That might seem hard to do, even after we are given my nuclear models, but to solve this we have to look at the way Methane is actually created. It isn't created *up* from Carbon, it is created *down* from Carbon Dioxide. Neither in nature nor in the lab do we see normal, unbonded Carbon in any sort of Hydrogen-rich environment becoming Methane. So in fact, Hydrogen *doesn't* want to create this covalent bond with Carbon, which sort of blows the whole mainstream analysis. In other words, it isn't electronegativities or other numbers that bring Hydrogen and Carbon together, which means all the hybridizing and merging was just smoke. We know it doesn't happen that way, so we should try to match the theory to the actual process, right? Let us look at the way nature creates Methane:



Therefore, to figure out how Methane creates the tetrahedron, we have to start with Carbon Dioxide and follow the process.



That is the mainstream diagram of Carbon Dioxide, which happens to match my nuclear diagrams precisely. According to my configurations, Carbon and Oxygen should bond in a line, and we know they do. Those bonds simply follow the main charge stream through the nuclei, and these nuclei are bonding pole to pole to pole.



[Sorry, the pink things are just photoshop ghosts.] The blue disks are alphas, the black disks are single protons. The "double covalent bonds" of CO2 are not created by electrons, as you see. They are created by protons aligning in the charge channels at the nuclear poles. In this way, nothing has to be

self-caused, or caused by the needs of theory. The protons do what they do not by monitoring one another or by jumping up to higher energies for no reason. They do what they do because the charge field causes them to do it. They are aligning to real streams of real charge photons which are moving all the time through the nucleus and between all nuclei. This immediately solves most of the problems of electron orbital theory. Mainstream theory is completely lacking in field mechanics, being nothing but pushes to match data. But once we have a real charge field to work with, we can explain these things mechanically. Most of the mystery evaporates with that one discovery.

That would be the diagram of CO<sub>2</sub> in a weak, balanced charge field. But we have seen that the charge field here on Earth is not balanced. It is also not weak. We think of STP (standard temperature and pressure) as normal, but it is actually quite high. That is why we see many molecules in a gaseous state, as with CO<sub>2</sub>. And the fact that the field is not balanced is indicated by many things. It is why we have [asymmetry in things like beta decay](#), as just one example. Which simply means that charge will be moving more strongly in one direction here than the other. If we sum the total field, we will find charge either moving left to right or right to left. Let us say it is moving left to right. Will that affect our diagram, and if so how?

Well, we have also found that [larger elements can re-arrange the protons of smaller elements](#), but only when they are brought nucleus-to-nucleus like this. The charge channels of Oxygen are stronger than the charge channels of Carbon, so when a Carbon nucleus is brought very close to an Oxygen nucleus, the Oxygen nucleus trumps the ambient charge field. In this case, Carbon is forced to channel the charge field coming out of Oxygen rather than the ambient field. In this way, Oxygen can “break” Carbon, forcing it to take configurations it couldn't normally take.

Since each blue disk here represents an alpha (two protons), each Oxygen *wants* to channel more charge than it is channeling. What I mean is that the core of Oxygen is composed of three alphas in parallel. These three alphas can channel more charge than is being pulled in by the single protons at the poles. In fact, that is why Oxygen is so reactive. The nuclear composition of Oxygen gives it a charge potential above its own charge strength. It *can* channel more than it normally *does* channel, taken alone. Therefore, it will tend to move into molecular situations where it can channel more. Pulling in Carbon is one way it can do this. Once we build the chain, each Oxygen in the chain is channeling more charge than it was alone. The core of the Oxygen nucleus is thereby closer to its channeling potential. And so, in a charge-rich environment like we see on Earth, the diagram of CO<sub>2</sub> will be this:



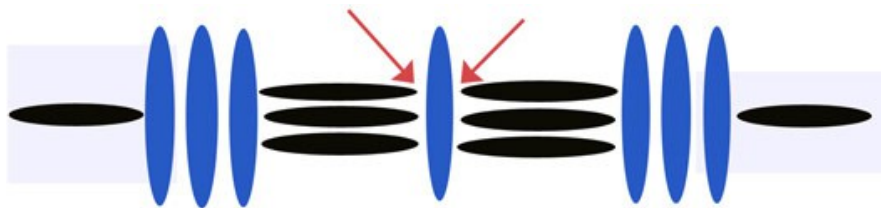
The Oxygens on each end have re-aligned the protons in the nucleus of Carbon, making it better able to channel charge along the main line. You see, if Carbon remained as before, its core alphas would be releasing charge radially instead of along the main line from left to right here. Any charge channeled out radially doesn't make it to the other Oxygen, and the power of the molecule is diminished. But if we have enough charge moving along the poles here (again, left to right), that charge stream can turn one of the core alphas of Carbon. Once that alpha is turned, the two protons in it can be used to

channel along the line, instead of releasing to the sides.

I assume that both core alphas are not being turned, since if we turn them both, there is nothing to focus the charge along the line at the center of the molecule. We need at least one core alpha diagrammed up and down here to channel the charge from left to right. Remember, the charge is moving through the poles of that alpha. It is *through* charge. But if we turn it sideways as well, then the center of our molecule loses its main vortex. There is nothing to prevent all those black disks from flying out of there.

You will say, “The strong charge stream should prevent them from flying out of there.” And given a strong enough charge stream, that may be true. There may be cases where both core alphas of Carbon do turn. However, in natural situations like we are looking at, I don't think that happens. It might be forced to happen in a lab, but it doesn't happen in nature.

Now, what happens if we insert Hydrogen into this equation? Well, we don't just “insert” Hydrogen, do we? The Hydrogen, being a gas, has a high velocity. And we can see that Carbon Dioxide is vulnerable to side traffic at its linkages. These are molecular linkages, remember, not fused linkages. If a proton comes in at high speed from the top or bottom of this last diagram, and happens to intersect at the black linkages (which are also protons), it can break the link. This will cause Carbon to be broken out of the CO<sub>2</sub>. Notice there are several ways the Carbon could break out of that molecule. At first glance, you would think the “normal” way would be to break out with two prongs on each end, leaving the third proton to go with Oxygen. That would give us a free Carbon of a peculiar nuclear configuration, but it wouldn't allow us to create Methane. But if we study the real collision of the incoming Hydrogen, we realize that isn't the way it would happen. The Hydrogen either has to hit a linking proton left or right:



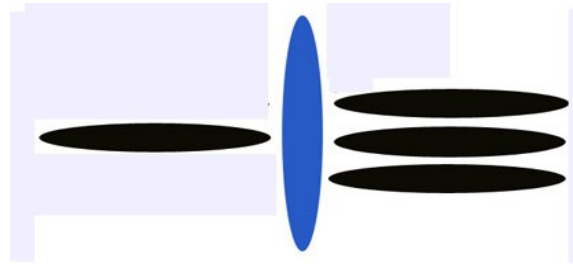
I have drawn the arrows in those places for a reason. That is the weak link in this chain. There is actually too much charge moving through that alpha. I have shown you in previous papers that an alpha can only take two proton's worth of charge moving through it. Here we have three. That configuration would not be stable on its own. It is stable here for two reasons: 1) it is a molecular configuration, not a fused configuration. Two of these protons aren't “plugged” into the holes. They are aligned, but not plugged. This allows some charge to dissipate into the ambient field. 2) this Carbon nucleus is not free: it is constrained on each side by a larger nucleus. In this position, most of the charge is not being fully recycled through the alpha, since by recycling we mean it is going in the poles and out the equators. Here, most of the charge is *through* charge, going straight through from pole to pole.

So although that configuration is stable—in the sense it is not prone to *spontaneous* dissolution—it is nonetheless weak. If the molecule is being hit from the sides by free protons, then that is the first place we would look for a break.

In fact, what will happen is that the incoming Hydrogens will be successfully bounced from all points

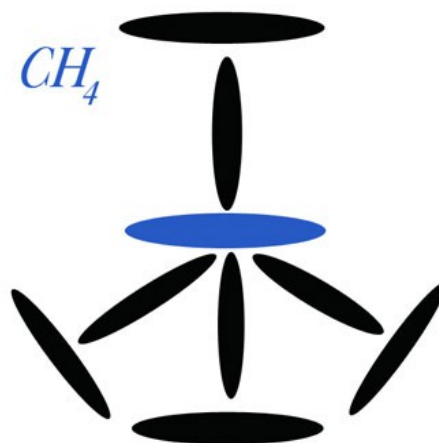
except those points. Only if a Hydrogen happens to hit there will it have a chance of causing a break-up of the molecule. And if a Hydrogen hits there and causes a break-up, *it may be captured*.

Here is why. If a Hydrogen hits there and causes a break-up of the molecule, the Oxygen on that end will escape, taking with it one proton on that side. That leaves two to go with the Carbon on that side, as you see. But since the Carbon just had three on that side a moment earlier, it will still have a charge stream set up for three. And so, the incoming Hydrogen—which *is* a proton—may be able to fall into that waiting slot. It isn't guaranteed to do so, but it *may* do so. If it does, the Carbon then has three protons on that side, so it only needs one on the other side to make up the requisite six. It then looks like this:



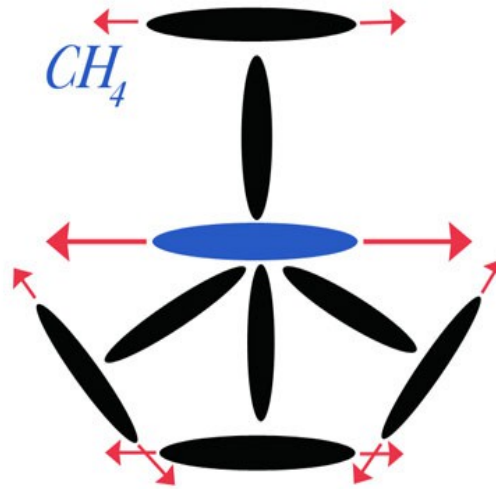
Although our normal Carbon nucleus wasn't prone to join to free Hydrogen, this nucleus broken by Oxygen is. Normal Carbon channels too strongly to allow Hydrogen to come in and bond. This is why Methane isn't made by bonding Hydrogen onto free Carbon. The Carbon nucleus first has to be broken and re-arranged by a larger nucleus. Only then can Hydrogen bond to it.

Why does this nucleus allow it? Because it only has one alpha at the core. Because of that alone, all the channels and vortices have been weakened. It acted as a strong through channeler for CO<sub>2</sub>, but once alone, it doesn't channel as strongly as the unbroken Carbon. It was a good through channeler in CO<sub>2</sub> only because it had Oxygen pushing charge in. But it no longer has that. It has to rely on the core alpha to pull charge in, and one blue disk obviously can't do that as well as two. This weakened Carbon nucleus is immediately vulnerable in a Hydrogen-rich environment, and Hydrogens will jump in to cap all four prongs, like so:

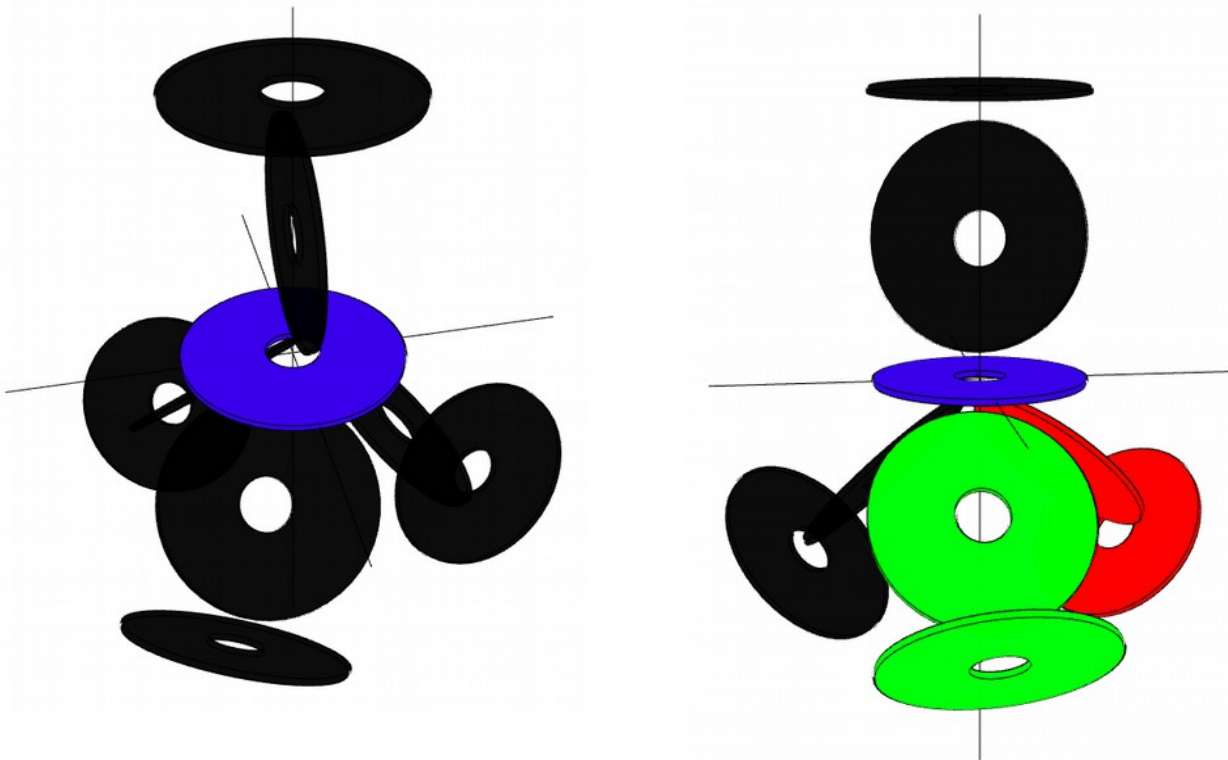


Carbon in its original two-alpha state wouldn't allow Hydrogen to get near enough to create those

molecular bonds, but the weakened configuration above does. As you see, this explains why the lower channels spread out. They aren't just filling space or dividing the three dimensions by four. Since we have charge channels that are channeling real photons, we have real exclusion. Those three cap protons on the bottom are not only excluding one another via their emitted charge streams, they are being excluded by the charge stream of the blue central alpha.



I have had my friend Arlo Emerson render these in 3D, to help you visualize this:



Once again, the blue disk is an alpha, the black disks are protons. In the second diagram, he has given each leg of the tripod a different color, so that you can differentiate them at a glance.



This means that the old theory was wrong in yet another way. It isn't the four Hydrogens that are excluding one another, as you see in the previous diagram with arrows. The top Hydrogen doesn't even come into the problem here. Once we have the right diagram, we can see that the top exclusion is caused by the core of the Carbon, not by the fourth Hydrogen. That blue disk is emitting charge out radially in a circle. So the three bottom Hydrogens are constrained by that charge stream as well as their own. We get the same angles either way, but in my diagram you can see exactly what is causing the angles. We have real field potentials here, not just a field divided by four for no mechanical reason.

This also means that the four Hydrogen positions are not strictly equivalent. If we follow the *amount* of charge being channeled, the top Hydrogen must be different than the other three. If we assume that the Methane will align to the ambient field with the three on the strong side\*, then the single Hydrogen on top will be channeling at  $\frac{1}{2}$ , while the others are channeling at  $\frac{1}{3}$ .

Another difference should be discovered in a strong applied magnetic field. Only one of the four Hydrogen links will align to the field, and that is the one residing on the pole. The other three will be at an angle. In fact, by following this Hydrogen, it is easy to discover the nuclear pole of the central Carbon.

This new theory allows me to make several other predictions, ones that can be checked with current machines—as far as I know. I believe it should be possible to measure the charge being emitted laterally by the Carbon inside CO<sub>2</sub>, and I think it will be found to be different from the lateral emission of Carbon alone. In other words, we need CO<sub>2</sub> in a solid (but still non-magnetic) state, and then we need to probe the Carbon in that molecule, mapping its lateral charge profile by some means. I believe we will find that the Carbon has a different nuclear make-up than Carbon alone. I think we will find that a single atom of Carbon has two alphas in its core, while an atom of Carbon in most molecules has one.

This ties into recent questions I have been asked about Fullerenes and irradiated graphites. It has been found that Carbon, although normally non-magnetic, can be very magnetic in some situations. I would suggest that the varying nuclear make-up of different forms of Carbon explains this in the most direct and mechanical way. It would appear that Carbon in compound with itself can re-arrange in the same way we saw it re-arranging in CO<sub>2</sub>, especially in an irradiated field or in long chains. Once you have two prongs on each end of Carbon and only one alpha in the core, this will create a spun-up through charge, which is what causes magnetism.

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I think you can now see that [I have once again destroyed electron orbital theory](#), in all its forms: MOT, VSEPR, etc. It is not only unnecessary and very ugly, it has almost no explanatory power. Even after all the non-mechanical pushes, it answers very few fundamental questions. Students have to be continually forbidden from asking the questions begged. But with my mechanical theory, we can answer all questions that arise, provided we are willing to do a little work and stay honest.

To see what I mean, we can return to the mainstream paper I was analyzing to start us off. In an inset note, Clark says,

People sometimes worry that the promoted electron is drawn as an up-arrow, whereas it started as a down-arrow. The reason for this is actually fairly complicated - well beyond the level we are working at. Just get in the habit of writing it like this because it makes the diagrams look tidy!

I find the tone of this writing really offensive and irritating. Students of chemistry and physics are now treated like first-graders—as if they can't comprehend a real physical theory. Even 4<sup>th</sup> year graduate students are talked down to in this manner. “Shut up and calculate!” “Just memorize the diagrams!” “We can't explain it to you at this level. It is a quantum mechanical problem and we know the answer at that level.”

But they don't. They can't explain this at any level, using current theory. You don't ever reach a level in physics or chemistry where they stop fudging you and start speaking sense. The quantum mechanical explanations just have a longer line of fudges than the fudges we have seen here. They don't become more mechanical, they become less mechanical. They don't become more sensible, they become less sensible. And they never stop treating you like a mischievous child, one that needs to be kept in line with psychological tricks. Richard Feynman was the grandmaster of these psychological tricks, but all the top physicists and chemists of the 20<sup>th</sup> century depended on them.

Just ask yourself this: If they really had a better explanation of any of this in quantum mechanics, why would they continue to teach this embarrassing line of pushes at lower levels? We can already see that the electron orbital theory doesn't work when it is stripped down to the basics, so how could it possibly work when it is tarted up with a bunch of fancy math? It can't. There is no way to correct fundamental theoretical contradictions and holes by applying complex equations to them. If electron orbital theory contradicts its own field definitions in the gloss, it is going to continue contradicting them in the fuller theory. That is, unless quantum mechanics corrects these field definitions. But of course it doesn't. If the fuller theory corrected the field definitions, they would not teach the contradictory field definitions in the gloss, would they? In short, if they had a good answer to any of these questions, they wouldn't teach this mainstream garbage at any level.

Besides which, I have just proven you can solve these problems with straightforward mechanics, and do it in a language that intelligent students at any level can understand. Smart highschool students can read this paper and follow it, even though as a matter of field theory it goes far beyond even what they teach in advanced quantum mechanics. It isn't littered with a lot of pushed math, and it never attempts to brainwash, propagandize, or finesse the reader. So you can be sure that physics and chemistry don't have to be in the form you have been sold. You don't have to accept what they are giving you, and I would recommend you refuse delivery from here on out. If your teachers start forcing electron orbital theory down your throat, ask them the hard questions. And when they start browbeating you after that—as they will—tell them to go get stuffed. Remember, “question authority” isn't just a T-shirt. It is the first order of science.

\*By strong side, I mean the photon side rather than the antiphoton side.